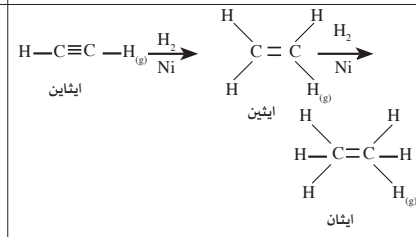
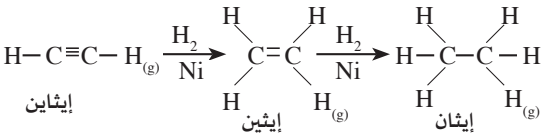
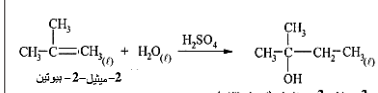
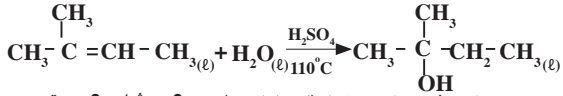
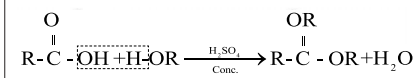
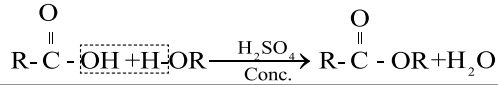
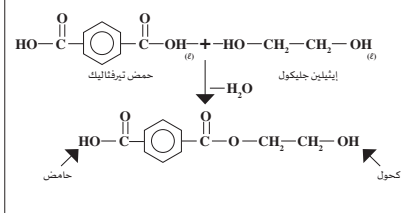
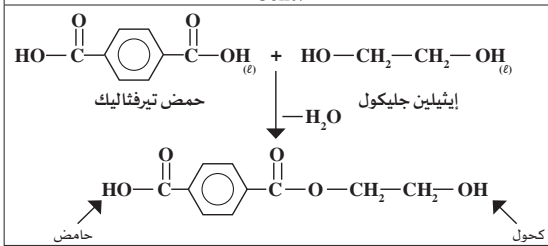
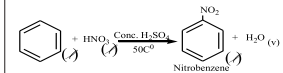
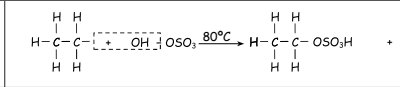
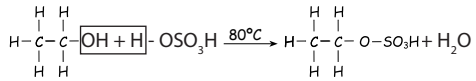
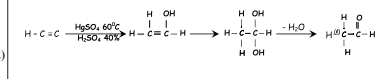
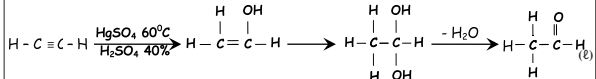
التصويبات الخاصة بالكيمياء طبقًا لأرقام الصفحات بكتاب المفاهيم علمى علوم

الصفحة	الخطأ	التصويب
133	$2FeO_{3(s)} + ½\ O_{2(g)} \xrightarrow{\Delta} Fe_2O_{3(s)}$ (حديد 69.6٪)	$2FeO_{(s)} + ½\ O_{2(g)} \xrightarrow{\Delta} Fe_2O_{3(s)}$ (حديد 69.6٪)
133	$2Fe_2O_3 3H_2O_{(s)} \xrightarrow{\Delta} 2Fe_2O_3 + 3H_2O_{(2)}$ (حديد 40٪) (حديد 69.6٪)	$2Fe_2O_3.3H_2O_{(s)} \xrightarrow{\Delta} 2Fe_2O_{3(s)} + 3H_2O_{(v)}$ (حديد 40٪) (حديد 69.6٪)
133	$4P_{(s)} + 5O_{2(g)} \xrightarrow{\Delta} 2PO_{5(g)}$	$4P_{(s)} + 5O_{2(g)} \xrightarrow{\Delta} 2P_2O_{5(s)}$
133	$3CO_{(g)} + Fe_2O_{3(s)} \xrightarrow{\Delta} 2Fe_{(s)} + 3CO_{2(s)}$	$3CO_{(g)} + Fe_2O_{3(s)} \xrightarrow[700^{\circ}C]{\text{أعلى من}} 2Fe_{(S)} + 3CO_{2(g)}$
133	$2CH_{4(s)} + CO_{2(g)} \xrightarrow{\Delta} 3CO_{(g)} + 5H_{2(g)}$	$2CH_{4(g)} + CO_{2(g)} + H_2O_{(v)} \xrightarrow{\Delta} 3CO_{(g)} + 5H_{2(g)}$
133	$2Fe_2O_{3(s)} + 3CO_{(s)} + 3H_{2(s)} \xrightarrow{\Delta} 4Fe_{(s)} + 3CO_{2(s)} + 3H_2O_{(s)}$	$2Fe_2O_{3(s)} + 3CO_{(g)} + 3H_{2(g)} \xrightarrow{\Delta} 4Fe_{(s)} + 3CO_{2(g)} + 3H_2O_{(v)}$
134	$3Fe_{(s)} + 4H_2O_{(s)} \xrightarrow{500^{5}C} Fe_3O_{4(s)} + 4H_{2(s)}$	$3Fe_{(s)} + 4H_2O_{(v)} \xrightarrow{500^{\circ}C} Fe_3O_{4(s)} + 4H_{2(g)}$
134	$2Fe_{(s)} + 3Cl_{2(s)} \xrightarrow{\Delta} 2Fe_3Cl_{3(s)}$	$2Fe_{(s)} + 3Cl_{2(g)} \xrightarrow{\Delta} 2FeCl_{3(s)}$
135	$Fe_{(s)} + H_2SO_{2(g)} \xrightarrow{dil} FeSO_{(aq)} + H_{2(g)}$	$Fe_{(s)} + H_2SO_{4(aq)} \xrightarrow{dil} FeSO_{4(aq)} + H_{2(g)}$
135	$Fe_{(s)} + 2HCl_{(aq)} \xrightarrow{dil} FeCl_{2(s)} + H_{2(s)}$	$Fe_{(s)} + 2HCl_{(aq)} \xrightarrow{dil} FeCl_{2(aq)} + H_{2(g)}$
135	$3Fe_{(s)} + 8H_2SO_{4(\ell)} \xrightarrow[conc]{} Fe_3SO_{4(aq)} + Fe_2(SO_4)_{2(aq)} + 4SO_{2(g)} + 8H_2O_{(s)}$	$3Fe_{(s)} + 8H_2SO_{4(\ell)} \xrightarrow[conc]{\Delta} FeSO_{4(aq)} + Fe_2(SO_4)_{3(aq)} + 4SO_{2(g)} + 8H_2O_{(v)}$
135	$\begin{matrix}COO\\ \\COO\end{matrix}Fe \xrightarrow[\text{معزل عن الهواء}]{\Delta} FeO_{(s)} + CO_{(g)} + CO_{2(g)}$	$\begin{matrix}COO\\ \\COO_{(s)}\end{matrix}Fe \xrightarrow[\text{معزل عن الهواء}]{\Delta} FeO_{(s)} + CO_{(g)} + CO_{2(g)}$
135	$Fe_2O_{3(s)} + H_{2(s)} \xrightarrow{400/700^{5}C} 2FeO_{(s)} + H_2O_{(s)}$	$Fe_2O_{3(s)} + H_{2(g)} \xrightarrow{400/700^{\circ}C} 2FeO_{(s)} + H_2O_{(v)}$
135	$Fe_3O_{4(s)} + H_{2(s)} \xrightarrow{400/700^{5}C} 3FeO_{(s)} + H_2O_{(s)}$	$Fe_3O_{4(s)} + H_{2(g)} \xrightarrow{400/700^{\circ}C} 3FeO_{(s)} + H_2O_{(v)}$
135	$4FeO_{(s)} + O_{2(g)} \xrightarrow{\Delta} 2Fe_3O_{3(s)}$	$4FeO_{(s)} + O_{2(g)} \xrightarrow{\Delta} 2Fe_2O_{3(s)}$
135	$Fe_{(s)} + H_2SO_{4(aq)} \xrightarrow{dil} FeSO_{4(aq)} + H_2O_{(\ell)}$	$FeO_{(s)} + H_2SO_{4(aq)} \xrightarrow{dil} FeSO_{4(aq)} + H_2O_{(\ell)}$
135	$2Fe(OH)_{3(s)} \xrightarrow{700^{\circ}C\ \text{أعلى من}} Fe_2O_{3(s)} + 3H_2O_{(g)}$	$2Fe(OH)_{3(s)} \xrightarrow{200^{\circ}C\ \text{أعلى من}} Fe_2O_{3(s)} + 3H_2O_{(v)}$
135	$2FeSO_{4(s)} \xrightarrow{\Delta} Fe_2O_{3(s)} + SO_{2(g)} + SO_{2(g)}$	$2FeSO_{4(s)} \xrightarrow{\Delta} Fe_2O_{3(s)} + SO_{2(g)} + SO_{3(g)}$
135	$Fe_2O_{3(s)} + 3H_2SO_{4(aq)} \xrightarrow[conc]{\Delta} Fe_2(SO_4)_{3(aq)} + 3H_2O_{(v)}$	$Fe_2O_{3(s)} + 3H_2SO_{4(\ell)} \xrightarrow[conc]{\Delta} Fe_2(SO_4)_{3(aq)} + 3H_2O_{(v)}$
135	$Fe_3O_{4(s)} + 4H_2SO_{4(\ell)} \xrightarrow[conc]{\Delta} FeSO_{4(aq)} + Fe_2(SO_4)_{3(aq)} + 4H_2O_{(s)}$	$Fe_3O_{4(s)} + 4H_2SO_{4(\ell)} \xrightarrow[conc]{\Delta} FeSO_{4(aq)} + Fe_2(SO_4)_{3(aq)} + 4H_2O_{(v)}$
135	$2Fe_3O_{4(s)} + ½O_{2(g)} \xrightarrow{\Delta} 3Fe_2O_{3(g)}$	$2Fe_3O_{4(s)} + ½O_{2(g)} \xrightarrow{\Delta} 3Fe_2O_{3(s)}$

136	2- التحليل (الكيفي) Quantitative Analysis	2- التحليل (الكمي) Quantitative Analysis
138	أبخرة اليود تظهر بلونها	أبخرة اليود تظهر بلونها
139	$FeSO_{4(aq)} + NO_{(g)} \longrightarrow FeSO_4.NO_{(s)}$	$FeSO_{4(aq)} + NO_{(g)} \longrightarrow FeSO_4.NO_{(s)}$
139	محلول ملح النحاس II $CuSO_{4(aq)} + H_2S_{(g)} \longrightarrow H_2SO_{4(al)} + CuS_{(s)}$	محلول ملح النحاس II $CuSO_{4(aq)} + H_2S_{(g)} \longrightarrow H_2SO_{4(aq)} + CuS_{(s)}$
145	عصير العنب 3.5 عصير الطماطم 4	عصير العنب 4 عصير الطماطم 4.2
146	$PbBr_{2(aq)}$ (آخر الصفحة)	$PbBr_{2(s)}$
147	$Cr \rightleftharpoons Cr^{2+} + 3e^-$ (فى المتسلسلة)	$Cr \rightleftharpoons Cr^{2+} + 2e^-$
149	$Fe(OH)_{2(s)}$ (المعادلة 3)	$2Fe(OH)_{2(s)}$
153	C_aH_{2n-2} C_nH_{2n} C_aH_{2n+2}	C_nH_{2n-2} C_nH_{2n} C_nH_{2n+2}
158		
164		
170		
173		
215		
250	molecular	moléculaire
250	Ka de $CH_3COOH = 1{,}8 \times 10^{-4}$	Ka de $CH_3COOH = 1{,}8 \times 10^{-5}$
253	Hydroxyde de baryum $Ba(OH)_2 \longrightarrow K^+$	تحذف
265		
269		
271